

# SAULT COLLEGE OF APPLIED ARTS AND TECHNOLOGY

SAULT STE. MARIE, ONTARIO



SAULT  
COLLEGE

## COURSE OUTLINE

**COURSE TITLE:** GENERAL CHEMISTRY

**CODE NO. :** CHM 107 **SEMESTER:** II

**PROGRAM:** PRE-HEALTH SCIENCES

**AUTHOR:** DAVID TROWBRIDGE

**DATE:** Jan 2011 **PREVIOUS OUTLINE DATED:** JAN 2010

**APPROVED:**

"B.Punch"

---

**CHAIR** **DATE**

**TOTAL CREDITS:** 5

**PREREQUISITE(S):** CHM 102

**HOURS/WEEK:** 3 HOURS LECTURE, 2 HOURS LAB

**Copyright ©2010 The Sault College of Applied Arts & Technology**  
*Reproduction of this document by any means, in whole or in part, without prior  
written permission of Sault College of Applied Arts & Technology is prohibited.*  
*For additional information, please contact Brian Punch, Chair,*  
*School of Natural Environment/Outdoor Studies & Technology Programs*  
*(705) 759-2554, Ext. 2681*

**I. COURSE DESCRIPTION:**

General Chemistry-II begins with an introduction to organic chemistry including nomenclature, and chemical reactions. Other main topics include chemical calculations, the mole concept, solutions chemistry, energy changes in chemical reactions, electrochemistry, oxidation-reduction equations, equilibrium in gaseous and aqueous reactions, acid/base equilibria and pH.

A comprehensive workshop on lab techniques and lab safety and on lab report preparation will be held during the early weeks of the semester in the laboratory class.

**II. LEARNING OUTCOMES:**

Upon successful completion of this course the student will demonstrate the ability to:

1. Describe the characteristics and apply the nomenclature of organic compounds.
2. Name chemical substances by common name and IUPAC name.
3. Describe the theory of ions in solutions, recognize precipitation, acid-base and gas forming reactions and write ionic and non ionic equations.
4. Explain the mole concept and quantify substances in terms of mass and moles and complete calculations to determine chemical formulas and quantities of substances involved in chemical reactions.
5. Describe and calculate energy changes in chemical reactions
6. Perform calculations involving compounds in aqueous solutions.
7. Write and balance oxidation – reduction equations.
8. Use equilibrium concepts to solve for desired quantities in gaseous and aqueous reactions
9. Apply the concepts of solubility product to solve problems in solubility equilibria.
10. Apply the concepts of acid – base theory to the solution of acid base equilibrium problems.

**III. ELEMENTS OF THE PERFORMANCE**

Upon successful completion of this course the student will demonstrate the ability to:

- 1) Describe the characteristics of organic compounds and name simple compounds.

## Potential Elements of the Performance:

- Identify the main functional groups that are common in organic compounds.
  - Name simple organic compounds when given their formula and write names given their name.
  - Predict and write common reactions of organic compounds
- 2) Perform calculations involving compounds in aqueous solutions.

## Potential Elements of the Performance:

- Make calculations for preparation of solutions having concentration expressed in moles per Litre or molarity (M), and ppm, and be able to convert from one concentration to another.
  - Using solubility rules, decide whether two soluble ionic compounds will or will not react to form a precipitate. If they will, write the net ionic equation.
  - Write the molecular equation, and then the net ionic equation for the neutralization of an acid and a base.
- 3) Write and balance oxidation – reduction reactions.

## Potential Elements of the Performance:

- Write a skeleton equation given as balanced oxidation-reduction equation. Label the oxidizing and reducing agents, the oxidized and reduced species, and the oxidation and reduction parts of the equation. Comment on the reaction by referring to the commonly observed oxidation states.
  - Given an oxidation-reduction equation (an unbalanced or a skeleton equation), complete and balance it by the half-reaction method and/or the oxidation number method.
- 4) Use equilibrium concepts to solve for the desired quantities in gaseous reactions.

## Potential Elements of the Performance:

- Given the starting amounts of reactants and the amount of one substance at equilibrium, find the equilibrium composition.
- Given the chemical equation, write the equilibrium-constant expression.

- Given the equilibrium composition, find  $K_c$ .
  - Given the concentrations of substances in a reaction mixture, predict the direction of reaction.
  - Given  $K_c$  and all concentrations of substances but one in an equilibrium mixture, calculate the concentration of this one substance.
  - Given the starting composition and  $K_c$  of a reaction mixture calculate the equilibrium composition.
  - Given a reaction, use Le Chatelier's principle to decide the effect of adding or removing a substance, changing the pressure, or changing the temperature.
- 5) Apply concepts of the solubility product to solve problems in solubility equilibria.

Potential Elements of the Performance:

- Given equilibrium concentrations calculate  $K_{sp}$  for a system in equilibrium
  - Given  $K_{sp}$  calculate equilibrium concentrations for simple systems
- 6) Apply the concepts of acid-base theory to the solution of acid-base equilibrium problems.

Potential Elements of the Performance:

- Calculate the pH during the titration of a strong acid and strong base, given the volumes and concentrations of the acid and base.
- Given the concentration of hydroxide ion (or concentration of strong base), calculate the hydrogen-ion concentration.
- Given the hydrogen ion concentration (or concentration of strong acid), calculate the pH; given the pH, calculate the hydrogen-ion concentration.
- Given the relative strengths of acids (or bases), decide whether reactants or products are favored at equilibrium.
- Decide whether an aqueous solution of a given salt will be acidic, basic, or neutral.

**IV. TOPICS**

1. Organic Chemistry Introduction
2. Chemical Reactions
3. Calculations with Chemical Formulas and Equations
4. Energy Changes in Reactions
5. Oxidation and Reduction Concepts
6. Equilibrium Concepts
7. Acid /Base Equilibria

**LABORATORY WORK**

In a laboratory setting, the student will conduct experimental procedures to support the theoretical concepts and these will be selected from the following and posted on LMS:

1. Construct models to represent the main types of organic compounds.
2. Investigate the physical and chemical properties of organic compounds.
3. Conduct chemical reactions and identify the products formed from the given reactants.
4. Determine the chemical formula of a compound formed in a chemical reaction based on mass and moles.
5. Recover a mass of a substance which has been subjected to a sequence of chemical reactions.
6. Titration of acids and bases, standardization and determination of an unknown acid.
7. Gravimetric determination of a chloride.
8. Volumetric determination of an unknown chloride.
9. Spectrophotometric determination of a metal in aqueous solution.

**V. REQUIRED RESOURCES/TEXTS/MATERIALS:**

Introductory Chemistry, C. Corwin, Prentice Hall, 2010

Lab Materials: Lab Coat, Safety Glasses

**VI. EVALUATION PROCESS/GRADING SYSTEM**

The following semester grades will be assigned to students in postsecondary courses:

Grade	Definition	Grade Point Equivalent
A+	90 – 100%	4.00
A	>80% - <90 %	3.00
B	>70 % - <80%	2.00
C	>60% - <70%	1.00
D	>50 % - <60%	0.00
F (Fail)	<50%	
CR (Credit)	Credit for diploma requirements has been awarded.	
S	Satisfactory achievement in field /clinical placement or non-graded subject area.	
U	Unsatisfactory achievement in field/clinical placement or non-graded subject area.	
X	A temporary grade limited to situations with extenuating circumstances giving a student additional time to complete the requirements for a course.	
NR	Grade not reported to Registrar's office.	
W	Student has withdrawn from the course without academic penalty.	

The final grade is calculated by adding the test marks (60%) and the term marks (40%). The lab mark is the sum of all marks awarded for the analysis plus the written report for each of the experiments. The analysis is graded on accuracy and precision. Assignments and quizzes are part of the term mark. The theory mark is the sum of all tests.

Tests (3 at 20% each)	60 marks
Term work - Lab Work /Quizzes/Assignments	<u>40 marks</u>
	100 marks

Assignments are due on the date specified at the beginning of the class. Late assignments will not be accepted so it is critical that you submit as much of the assignment as possible on the due date. Lab reports are due one week from completion of the lab. Late labs will be downgraded 10% per day after agreed due date. (See details below regarding missed labs)

**ATTENDANCE:**

Your grade will be greatly affected by attendance at scheduled classes and labs. 85% is required at all theory classes while 100% is needed for all labs. Serious illness (doctor's medical slip) is the only valid excuse.

Each laboratory activity requires a pre-lab assignment in which the student will familiarize him/herself with the procedure, equipment and safety concerns. This will include researching the hazards and precautions for each chemical used in the lab as described in the Material Safety Data Sheets (MSDS) to be found in the chemistry lab.

A pre-lab lecture will discuss the theoretical and practical aspects of the lab and identify any safety precautions. Consequently attendance at these pre-lab lectures is mandatory and any student missing this lecture must prepare a step by step procedure which also must identify all safety concerns before you will be allowed to begin the lab.

Labs missed without prior agreement of the instructor will be give a **maximum of 50%** of the lab mark once completed to the instructor's satisfaction.

**VII. SPECIAL NOTES:****Course Outline Amendments:**

The professor reserves the right to change the information contained in this course outline depending on the needs of the learner and the availability of resources.

**Retention of Course Outlines:**

It is the responsibility of the student to retain all course outlines for possible future use in acquiring advanced standing at other postsecondary institutions.

**Prior Learning Assessment:**

Students who wish to apply for advance credit transfer (advanced standing) should obtain an Application for Advance Credit from the program coordinator (or the course coordinator regarding a general education transfer request) or academic assistant.

Students will be required to provide an unofficial transcript and course outline related to the course in question. Please refer to the Student Academic Calendar of Events for the deadline date by which application must be made for advance standing.

Credit for prior learning will also be given upon successful completion of a challenge exam or portfolio.

Substitute course information is available in the Registrar's office.

**Disability Services:**

If you are a student with a disability (e.g. physical limitations, visual impairments, hearing impairments, or learning disabilities), you are encouraged to discuss required accommodations with your professor and/or the Disability Services office. Visit Room E1101 or call Extension 2703 so that support services can be arranged for you.

**Communication:**

The College considers WebCT/LMS as the primary channel of communication for each course. Regularly checking this software platform is critical as it will keep you directly connected with faculty and current course information. Success in this course may be directly related to your willingness to take advantage of the Learning Management System communication tool.

**Plagiarism:**

Students should refer to the definition of “academic dishonesty” in Student Code of Conduct. A professor/instructor may assign a sanction as defined below, or make recommendations to the Academic Chair for disposition of the matter. The professor/instructor may (i) issue a verbal reprimand, (ii) make an assignment of a lower grade with explanation, (iii) require additional academic assignments and issue a lower grade upon completion to the maximum grade “C”, (iv) make an automatic assignment of a failing grade, (v) recommend to the Chair dismissal from the course with the assignment of a failing grade. In order to protect students from inadvertent plagiarism, to protect the copyright of the material referenced, and to credit the author of the material, it is the policy of the department to employ a documentation format for referencing source material.

**Student Portal:**

The Sault College portal allows you to view all your student information in one place. mysaultcollege gives you personalized access to online resources seven days a week from your home or school computer. Single log-in access allows you to see your personal and financial information, timetable, grades, records of achievement, unofficial transcript, and outstanding obligations. Announcements, news, the academic calendar of events, class cancellations, your learning management system (LMS), and much more are also accessible through the student portal. Go to <https://my.saultcollege.ca>.



**Electronic Devices in the Classroom:**

Students who wish to use electronic devices in the classroom will seek permission of the faculty member before proceeding to record instruction. With the exception of issues related to accommodations of disability, the decision to approve or refuse the request is the responsibility of the faculty member. Recorded classroom instruction will be used only for personal use and will not be used for any other purpose. Recorded classroom instruction will be destroyed at the end of the course. To ensure this, the student is required to return all copies of recorded material to the faculty member by the last day of class in the semester. Where the use of an electronic device has been approved, the student agrees that materials recorded are for his/her use only, are not for distribution, and are the sole property of the College.

**Attendance:**

Sault College is committed to student success. There is a direct correlation between academic performance and class attendance; therefore, for the benefit of all its constituents, all students are encouraged to attend all of their scheduled learning and evaluation sessions. This implies arriving on time and remaining for the duration of the scheduled session.

**Tuition Default:**

Students who have defaulted on the payment of tuition (tuition has not been paid in full, payments were not deferred or payment plan not honoured) as of the first week of March will be removed from placement and clinical activities. This may result in loss of mandatory hours or incomplete course work. Sault College will not be responsible for incomplete hours or outcomes that are not achieved or any other academic requirement not met as of the result of tuition default. Students are encouraged to communicate with Financial Services with regard to the status of their tuition prior to this deadline to ensure that their financial status does not interfere with academic progress.